

Concentration molaire :

$$C_M = (4.5/58.5) * (1/0.5) = 0.154 \text{ mol.L}^{-1}$$

$$C_M = 154 \text{ mM}$$

Concentration ionique :

$$C_i = \nu \cdot C_M \quad (\nu = \text{nombre de particules} = 2)$$

$$C_i = 2 \times 0,154 = 0,308 \text{ ion.L}^{-1}$$

Concentration osmolaire :

$$C_{\text{osm}} = \beta \cdot C_M$$

$$\beta = 1 + \alpha (\nu - 1) = 2 \quad [\alpha = \text{taux de dissociation}]$$

$$C_{\text{osm}} = 2 \cdot C_M = 2 \times 0,154 = 0,308 \text{ osmol.L}^{-1}$$